

THE DEPARTMENT OF CHEMISTRY AND ENVIRONMENTAL SCIENCE

Chemical Technology – Lecture & Lab  
Chem301-101  
Fall 2025 Course Syllabus

**NJIT Academic Integrity Code:** All Students should be aware that the Department of Chemistry & Environmental Science (CES) takes the University Code on Academic Integrity at NJIT very seriously and enforces it strictly. This means that there must not be any forms of plagiarism, i.e., copying of homework, class projects, or lab assignments, or any form of cheating in quizzes and exams. Under the University Code on Academic Integrity, students are obligated to report any such activities to the instructor.

**COURSE INFORMATION:** *Combination of Lecture and Laboratory components*

**Course Description:** Chem 301 - Chemical Technology. Introduction to chemistry geared for students in Engineering Technology. No prior knowledge of chemistry is assumed or required.

**Lecture:** Attendance is taken at the beginning of each lecture. You may miss one lecture without excuse. Beyond one absence evidence of an unavoidable situation (Doctors Note, Police Report of a car accident, Jury Duty summons, National Guard Orders, etc.) will be necessary to avoid loss of participation points for the class. All exams are given during Lecture Hours.

**Laboratory:** As Qualification of Engineering Programs at NJIT require laboratory experience **Attendance in LAB is MANDATORY:** If you cannot come to lab during your designated in-lab period please inform your instructor as soon as possible. You may be able to work with a different Lab Set for that lab. Missing 1 lab will not affect your grade but **missing 3** in person Lab classes for any reason will result in a **failing grade (F)**. **Missing more than one Lab for an excused reason (approved by DoS and/or OARS) will require taking the lab separately to make your degree qualification requirement.**

Number of Credits: 3

Prerequisites: none

Course-Section and Instructors

Course-Section	Day	Time	Location	Instructor
Chem 301-101	Wednesday	6:00 PM-7:55 PM	Tiernan 114	Dr. Xiaokai Zhang
	Friday	8:05 PM-10:00 PM	Tiernan 208	

Office Hours: by appointment (Email: [xz458@njit.edu](mailto:xz458@njit.edu) Cell: 201-628-4879)

Textbook: (Hard Copy or Electronic, your choice)

Title	Chemistry for Engineering
Author	Brown & Holme
Edition	3 <sup>rd</sup> or 4 <sup>th</sup>
Year	2015 and on
Publisher	Cengage
ISBN #	978-1-337-39890-9

**Required Lab manual:** Chem301 F2025 Lab Manual is available in the NJIT bookstore.

**Also Required:**

- Scientific calculator: capable of handling logs & exponentials. No cell phone, programmable or any other communications/internet active calculator will be permitted on exams.
- Personal Protective Equipment (PPE) – ALL Required:
  - Goggles: you must wear protective eye wear whenever you are in lab. If you wear glasses, purchase side shields or goggles designed to fit over your glasses.
  - Lab coat: an extra protective layer between your body and your experiment. Designed to cover and protect your arms and your body from your neck down to your knees. It also protects your clothes.
  - Shoes: you must wear closed shoes in lab. The shoe itself must cover your entire foot. Sox with sandals are not acceptable.
  - Gloves: Nitrile (not latex). A box of gloves is good for 5 people so consider a group purchase. If nitrile gloves are not what is required, we will provide the appropriate substitute.

**University-wide Withdrawal Date:** The last day to withdraw with a **W** is Monday, November 10, 2025. It will be strictly enforced.

**POLICIES**

**All CES students must familiarize themselves with, and adhere to, all official university-wide student policies.** CES takes these policies very seriously and enforces them strictly.

**Grading Policy:** The final grade in this course will be determined as follows:

**You must pass BOTH the lecture and the lab portions of the course to be eligible to pass the entire course**

**LECTURE (60%):**

Homework & Exams collectively = 1 exam grade		Maximum Points
Exam I - tentatively Oct-08	} Top 2 scores =	100
Exam II tentatively Nov-12		100
Exam III tentatively Dec-17		100
Maximum Test Points		200
Homeworks		60
Class participation (attendance, answering & asking questions)		40
<b>Total Grade Lecture Component</b>		<b>300</b>

- Exams are open book. You may use the textbook (hardcopy), course notes, homework assignments, and labs. You **may not** use any other resources, including the internet. All exams have a periodic table of the elements on the last page.
- Exams will either be On-line or in person as per university guidelines. In person exams are the rule with on-line exams given only in accordance with university restrictions.
- On-line exams are given via the Respondus lockdown browser. You must have a computer capable of running this browser to take the exam. The NJIT library has several computer stations running the Respondus lockdown browser if your computer cannot support this browser.
- **1 exam score must be at least 60 to pass the class. The instructors remind you that the first exam is the easiest of the 3 exams. If you feel yourself excessively challenged in Chemistry this is the exam to focus on to satisfy the 60 point requirement.**
- The only exam extra credit is on the exams.
- Homework is submitted online to the CANVAS portal and is due at the *beginning* of lecture & is gone over in class; NO late homework is accepted but the lowest score for one homework is dropped.
- Homework points are converted to percent at the end of the semester.

**Makeup Exam Policy:** Since 1 exam may be exchanged for the homework grade will normally be **NO MAKE-UP QUIZZES OR EXAMS** during the semester. In extraordinary instances due to extended illness or mandated quarantine arrangements may be made with your professor. If a student has a legitimate reason

for an extended illness or mandated quarantine, the student should contact the Dean of Students office and present written verifiable proof.

**LAB (40%):** Make sure you sign the attendance sheet after you have completed the lab, cleaned up and put back any instruments and glassware (cleaned) have been put away. There are 1 lab safety training and 10 labs in this course. The lowest grade will be dropped for a total of 10 marked labs. Each lab will be graded on the point scale as detailed below. Everyone submits every component of the lab. All submissions are online.

	Maximum points per lab
Pre-Lab Worksheet(s)	10
In-Lab: technique, adherence to safety procedures and lab	5
Lab Worksheet(s)	15
Total Grade Laboratory Component	25
8 Graded Labs Total	200 Points

- **Pre-lab worksheets must be submitted before the student is allowed to do In-Lab experimentation.** Pre-Lab worksheets are submitted *one day before the laboratory work* via the appropriate CANVAS Portal.
- **Anyone without a graded pre-lab worksheet or not wearing the appropriate PPE will not be allowed in lab – no exceptions.**
- In lab teams of 4-5 students perform the experiments. Each student must be an active participant in preparing and performing the lab.
- Although a team completes **each lab every team member submits her or his own completed lab sheet.** Obviously, team members sheets will contain the same data and much of the same analysis. NOTE: questions on lab techniques, data analysis, and conclusions have appeared on past exams. It pays to understand each lab so you can score these points on the exam.
- **You cannot submit data you did not acquire yourself with your team.** If you do not attend the lab, you cannot use the data that your regular lab team acquired, and you are missing that lab. You may be scored for the pre-lab if it is submitted before the deadline. Attempts to turn in a post-lab summary will be considered cheating. All lab members will receive a 0 for the lab (cannot be dropped). Due dates for each lab are clearly posted in CANVAS. No late lab submissions will be accepted.
- **Keep track of your lab grades: missing 3 labs regardless of the reason** will result in an F in the lab portion of the course and therefore an F grade for the course.

Your final letter grade in this course will be based on the following *tentative* point score out of 300 + 200 =500 possible points:

A	>440	C	320-350
B+	410-440	D	275-320
B	380-410	F	<275
C+	350-380		

The lecturers are empowered to raise grades by one half scale (*i.e.* C → C<sup>+</sup>) in instances where the point score to the next step is close. This will only be considered for individuals who have submitted **every** homework assignment. Do your homework... it pays off in more ways than one.

**Cellular Phones:** All cellular phones and other electronic devices must be switched off during all class times. During exams cell phones and smart watches will join your books in your backpack at the front of the classroom. During lab students can take *brief* calls by stepping out of the lab room. Habitual or lengthy interruptions will result in penalties.

## LEARNING OUTCOMES:

- Analyze problems using the scientific method.
- Make computations using metric system units & be able to convert between units.
- Factor in experimental limits in precision when doing calculations.
- Explain how atomic components and their arrangement dictate periodic trends.
- Form ionic compounds from constituent metals, nonmetals, and polyvalent ions.
- A basic understanding of covalent bonding.
- Apply conservation of mass to balancing chemical reactions.
- Determine empirical and molecular formulas
- Balance chemical equations.
- Determine quantities of reactants required or of products produced in a given chemical reaction using the principles of stoichiometry.
- Determine concentrations of aqueous solutions: molarity, mole fraction.
- Determine volumes or concentrations of reactants required or of products produced in a given aqueous reaction using the principles of solution stoichiometry.
- Determine unknown concentrations in acid-base titration reactions.
- Use concentration or density to convert between volume and mass.
- Analyze the effects of intermolecular forces on liquid systems in terms of their effects on physical properties including boiling points, vapor pressure, and solubility.
- Describe the flow of electrons in oxidation-reduction reactions.
- Analyze the conversion and transfer of energy or heat in a chemical reaction.
- Understand the difference between thermodynamic & kinetic effects.
- Analyze voltaic (Galvanic) cells
- Compute cell potentials.
- Understand the operation of batteries & fuel cells

## ADDITIONAL RESOURCES

**Chemistry Tutoring Center:** Located in the Central King Building, Lower Level, Rm. G12. Hours of operation are Monday – Friday 10:00 am - 6:00 pm. For further information please click [here](#).

### Good on-line general textbook sources:

1. Chem1 virtual chemistry textbook: <http://www.chem1.com/acad/webtext/virtualtextbook.html>
2. ChemMystery is geared for high school but a lot of it applies here too. <http://library.thinkquest.org/3659/>

**Accommodation of Disabilities:** Office of Accessibility Resources and Services (*formerly known as Disability Support Services*) offers long term and temporary accommodations for undergraduate, graduate and visiting students at NJIT.

If you need accommodations due to a disability, please contact Chantonette Lyles, Associate Director at the Office of Accessibility Resources and Services at **973-596-5417** or via email at [lyles@njit.edu](mailto:lyles@njit.edu). The office is in Fenster Hall Room 260. A Letter of Accommodation Eligibility from the Office of Accessibility Resources Services office authorizing your accommodations will be required.

For further information regarding self-identification, the submission of medical documentation and additional support services provided please visit the Accessibility Resources and Services (OARS) website at:

- <http://www5.njit.edu/studentssuccess/disability-support-services/>

<b>Sep</b>	<b>2</b>	<b>First Day of Classes</b>
<b>Sep</b>	<b>8</b>	<b>Last Day to Add/Drop a Class</b>
<b>Sep</b>	<b>8</b>	<b>Last Day for 100% Refund, Full or Partial Withdrawal</b>
<b>Sep</b>	<b>9</b>	<b>W Grades Posted for Course Withdrawals</b>
<b>Sep</b>	<b>15</b>	<b>Last Day for 90% Refund, Full or Partial Withdrawal, No Refund for Partial Withdrawal after this date</b>
<b>Sep</b>	<b>29</b>	<b>Last Day for 50% Refund, Full Withdrawal</b>
<b>Oct</b>	<b>20</b>	<b>Last Day for 25% Refund, Full Withdrawal</b>
<b>Nov</b>	<b>10</b>	<b>Last Day to Withdraw</b>
<b>Nov</b>	<b>26</b>	<b>Friday Class Meet</b>
<b>Nov</b>	<b>27</b>	<b>Thanksgiving Recess Begins</b>
<b>Nov</b>	<b>30</b>	<b>Thanksgiving Recess Ends</b>
<b>Dec</b>	<b>11</b>	<b>Last Day of Classes</b>
<b>Dec</b>	<b>14</b>	<b>Final Exams Begin</b>
<b>Dec</b>	<b>20</b>	<b>Final Exams End</b>
<b>Dec</b>	<b>22</b>	<b>Final Grades Due</b>

**Fall 2025 LECTURE, HOMEWORK AND LAB SCHEDULE.**

Wk	Wednesday	Lecture	Homework (25 pts) & <i>pre-LAB (10 pts) Due @ start of class</i>	Lab (in person), Lab worksheet (15 pts) due @ 11:59P Friday
1	Sep-03	Introduction to course.	N/A	No lab
2	Sep-10	Ch 01: Numbers & measurements, Density, units & unit conversions. Physical Properties, separation of mixtures.	Understanding Syllabus (5 pts), EXAM 1 HWK 1, <i>Pre-Lab 01: Measurement and Density</i>	Lab 01: Measurement & Density (in lab)
3	Sep-17	Ch 02: atomic structure & mass, Isotopes; periodic table & trends. Ions, ionization energy,	EXAM 1 HWK 2, <i>Pre-Lab 02: Atoms &amp; Ions</i>	Lab 02: Synthesis of Soap
4	Sep-24	Ch 07 sec 2: ionic compounds – criss cross rule. Ch 03: Chemical Formulas, hydrates, formula mass, % composition, moles	EXAM 1 HWK 3	Lab 03: Determination of the Chemical Formula of a copper hydrate
5	Oct-01	Empirical & molecular formulas, elemental analysis for formula determination, balancing eqns. Stoichiometry, theoretical yield, percent yield		Lab 04: Chemical Equations; Synthesis of Zinc Iodide.
6	Oct-08	<b>EXAM 1 (chapters 1, 2 &amp; part 3)</b>	EXAM 1 HWK 4, <i>Pre-Lab 3: Hydrate analysis</i>	No Lab
7	Oct-15	Limiting Reagents, solutions: units, dilution, solution chemistry	None	No Lab.
8	Oct-22	Acids & Bases, Acid-base titrations; Ch 05: Gases pt1	EXAM 2 HWK 1	Lab 5 Experiment: Acid-Base Titration & Reaction (Handout).
9	Oct-29	Ch 05: Gases pt2	EXAM 2 HWK 2, <i>Pre-Lab 4: Balancing Chemical Equations</i>	Lab 6 Calorimetry Experiment.
9	Nov-05	Ch 09: Energy Exchange and Thermodynamics	EXAM 2 HWK 3, & Concentration and molarity, <i>pre-Lab 5: Titration prelab</i>	No Lab.
10	Nov-12	<b>EXAM 2 (ch 3 – 5)</b>	EXAM 2 HWK 4, <i>Pre-Lab 6: Calorimetry Simulations</i>	Lab 7: Electrochemistry Experiment
11	Nov-19	Ch 13: Electrochemistry pt1	None	No Lab
12	Nov-26	Ch 13: Electrochemistry pt2	EXAM 3 HWK 1, <i>Pre-Lab 7: Electrochemistry</i>	Lab 8: Daniell cell
13	Dec-03	Electrochemistry review		No Lab
14	Dec-17	<b>EXAM 3 (ch 9 &amp; 13)</b>	EXAM-3 HWK 3	No Lab